

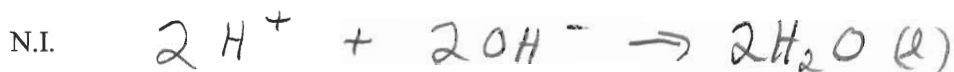
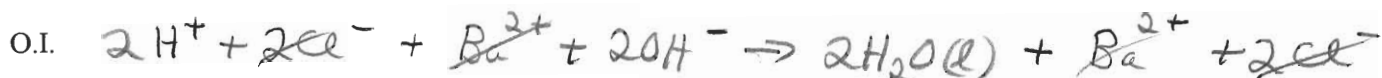
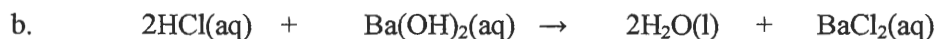
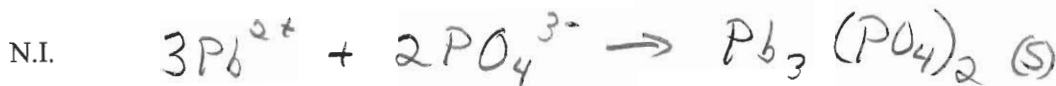
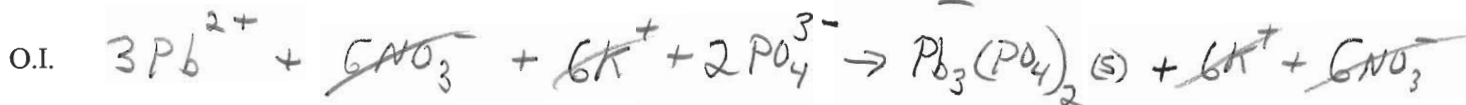
1. (4 Pts) Give the correct name AND formula for 2 strong acids and 2 strong bases.

Acids (Refer to notes or Book) Bases

2. (9 Pts) Complete the following table.

Element or ION	No. of protons	No. of electrons	No. neutrons
$^{15}\text{N}^{3-}$	7	$7+3=10$	$15-7=8$
^{55}Fe	26	26	$55-26=29$
^{37}Cl	17	17	$37-17=20$

3. (8 Pts) Give the overall ionic (O.I.) and the net ionic (N.I.) equation for the following reactions.



4. (4 Pts) A certain medication calls for a dosage of 1.0 μL for each kilogram of body weight. If a patient weighs 193 pounds, how many μL of medication should his dosage contain? Conversion Factors: $\mu = 10^{-6}$, $k = 10^3$, 1 kg = 2.205 pounds

$$\frac{193 \cancel{\text{lbs}}}{2.205 \cancel{\text{lbs}}} \times \frac{1 \cancel{\text{kg}}}{1 \text{kg}} \times \frac{1.0 \text{ mL}}{1 \cancel{\text{kg}}} = 87.5 \text{ or } \underline{\underline{88 \text{ mL}}}$$

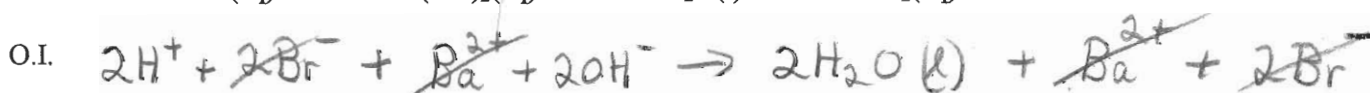
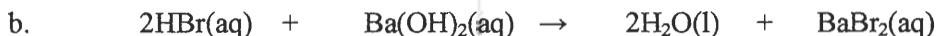
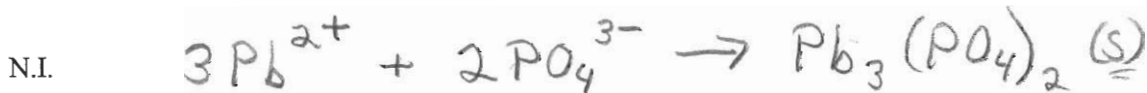
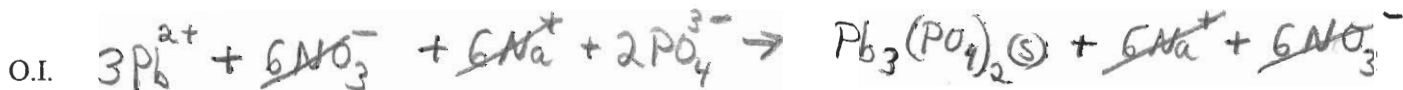
1. (4 Pts) Give the correct name AND formula for 2 strong acids and 2 strong bases.

Bases (Refer to notes or Book) Acids

2. (9 Pts) Complete the following table.

Element or ION	No. of protons	No. of electrons	No. neutrons
$^{16}\text{O}^{2-}$	8	10 (8+2)	8
^{56}Fe	26	26	56-26 = 30
^{35}Cl	17	17	35-17 = 18

3. (8 Pts) Give the overall ionic (O.I.) and the net ionic (N.I.) equation for the following reactions.



4. (4 Pts) A certain medication calls for a dosage of 2.0 μL for each kilogram of body weight. If a patient weighs 173 pounds, how many μL of medication should his dosage contain? Conversion Factors: $\mu = 10^{-6}$, $\text{k} = 10^3$, 1 kg = 2.205 pounds

$$\frac{173 \text{ lbs} \mid 1 \text{ kg}}{2.205 \text{ lbs} \mid 1 \text{ kg}} \times \frac{2.0 \mu\text{L}}{1 \text{ kg}} = \underline{157 \mu\text{L}}$$

160 μL