

Table 1. POLYATOMIC IONS FILL IN THE MISSING NAMES

-1 Anions	
Formula	Name
CN^-	
SCN^-	
OH^-	
NO_3^-	
NO_2^-	
CH_3COO^- (or $\text{C}_2\text{H}_3\text{O}_2^-$)	
MnO_4^-	
ClO_4^-	
ClO_3^-	
ClO_2^-	
ClO^- (or OCl^-)	
HCO_3^-	
HSO_4^-	

*The "oxy-anions" for the elements bromine and iodine are named in a manner analogous to that shown here for chlorine.

**The International Union of Pure and Applied Chemistry has recommended that use of the names *bicarbonate* and *bisulfate* be discontinued.

-2 Anions	
Formula	Name
CO_3^{2-}	
$\text{C}_2\text{O}_4^{2-}$	
SO_4^{2-}	
$\text{S}_2\text{O}_3^{2-}$	
SO_3^{2-}	
$\text{C}_8\text{H}_4\text{O}_4^{2-}$	
CrO_4^{2-}	
$\text{Cr}_2\text{O}_7^{2-}$	
O_2^{2-}	
-3 Anions	
PO_4^{3-}	
PO_3^{3-}	
AsO_4^{3-}	
+1 Cations	
NH_4^+	
H_3O^+	
+2 Cation	
Hg_2^{2+}	

**Table 1. POLYATOMIC IONS. FILL
IN THE MISSING FORMULA AND ITS
CHARGE**

-1 Anions	
Formula	Name
	cyanide
	thiocyanate
	hydroxide
	nitrate
	nitrite
	acetate
	permanganate
	perchlorate*
	chlorate*
	chlorite*
	hypochlorite*
	hydrogen carbonate (or bicarbonate**)
	hydrogen sulfate (or bisulfate**)

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-2 Anions	
Formula	Name
	carbonate
	oxalate
	sulfate
	thiosulfate
	sulfite
	phthalate
	chromate
	dichromate
	peroxide
-3 Anions	
	phosphate
	phosphite
	arsenate
+1 Cations	
	ammonium
	hydronium
+2 Cation	
	mercury(I)

