Ternary acids (those derived from polyatomic ions):

a. If the polyatomic ion ends is **ate**, change the **ate** to **ic** and add the word acid.

b. If the polyatomic ion ends is **ite**, change the **ite** to **ous** and add the word acid.

1. Acid Name **Sulfuric acid**	Polyatomic ion and its name SO_4^{2-} sulf**ate**	Acid formula $H_2SO_4(aq)$
2. Acid Name Phosphor**ic** acid	Polyatomic ion and its name PO_4^{3-} phosph**ate**	Acid formula **$H_3PO_4(aq)$**
3. Acid Name Sulfur**ous** acid	Polyatomic ion and its name **SO_3^{2-}** _sulf**ite**_____	Acid formula $H_2SO_3(aq)$
4. Acid Name **Nitric acid**	Polyatomic ion and its name NO_3^{1-} nitr**ate**	Acid formula $HNO_3(aq)$
5. Acid Name Oxal**ic** acid	Polyatomic ion and its name $C_2O_4^{2-}$ oxal**ate**	Acid formula **$H_2C_2O_4$ (aq)**
6. Acid Name Arsenic acid	Polyatomic ion And its name **AsO_4^{3-}** arsen**ate**	Acid formula $H_3ASO_4(aq)$
7. Acid Name **Phosphorous acid**	Polyatomic ion And its name PO_3^{3-} phosph**ite**	Acid formula $H_3PO_3(aq)$