

**Show All Work To Receive Credit!** Conversion factors and prefixes:

$G = 10^9$ ,  $M = 10^6$ ,  $k = 10^3$ ,  $c = 10^{-2}$ ,  $m = 10^{-3}$ ,  $\mu = 10^{-6}$ ,  $n = 10^{-9}$ ,  $2.54 \text{ cm} = 1 \text{ in}$ ,  
 $12 \text{ in} = 1 \text{ ft}$ ,  $5280 \text{ ft} = 1 \text{ mile}$ ,  $3 \text{ feet} = 1 \text{ yd}$ ,  $60 \text{ sec} = 1 \text{ min}$ ,  $1 \text{ hr} = 60 \text{ min}$ ,  $4 \text{ quarts} = 1 \text{ gal}$ ,  $2 \text{ pints} = 1 \text{ quart}$

1. (6 Pts) Perform each of the following conversions. You must show the complete setup.

a. Convert 65 mL to nL.

$$\frac{65 \text{ mL}}{1 \text{ mL}} \times \frac{10^{-3} \text{ L}}{1 \text{ L}} \times \frac{1 \text{ nL}}{10^{-9} \text{ L}} = 65 \times 10^6 \text{ nL} \quad (6.5 \times 10^7 \text{ nL})$$

b. Convert 85  $\mu\text{L}$  to mL.

$$\frac{85 \mu\text{L}}{1 \mu\text{L}} \times \frac{10^{-6} \text{ L}}{1 \text{ L}} \times \frac{1 \text{ mL}}{10^{-3} \text{ L}} = 85 \times 10^{-3} \text{ mL} \quad (8.5 \times 10^{-2} \text{ mL})$$

2. (4 Pts) Assume each of following numbers are measurements. Perform the indicated operations and then report the answer with the proper number of significant figures.

Limited to 0.01 place

a.  $2.14 \text{ cm} + 12.126 \text{ cm} + 0.12 \text{ cm} = 14.386 \text{ cm} \rightarrow 14.39 \text{ cm}$

b.  $1.25 \text{ cm} \times 2.1 \text{ cm} \times 1.145 \text{ cm} = 3.0 \text{ cm}^3$  (2 s.f.)

3. (5 Pts) A poster measures 22 inches by 44 inches. Determine its area in square cm.

$$\frac{22 \text{ in}}{1 \text{ in}} \times \frac{2.54 \text{ cm}}{1 \text{ in}} \times \frac{44 \text{ in}}{1 \text{ in}} \times \frac{2.54 \text{ cm}}{1 \text{ in}} = 6245 \text{ cm}^2 \quad (6200 \text{ cm}^2)$$

4. (5 Pts) How many inches are in 47 miles?

$$\frac{47 \text{ miles}}{1 \text{ mile}} \times \frac{5280 \text{ ft}}{1 \text{ mile}} \times \frac{12 \text{ in}}{1 \text{ ft}} = 2977920 \text{ in}$$

5. (5 Pts) A sample of silver ore was found to contain 0.56 % silver by mass. How many grams of silver can be recovered 800.0 kg of ore?

$$\frac{800.0 \text{ kg ore}}{1 \text{ kg}} \times \frac{10^3 \text{ g}}{1 \text{ kg}} \times \frac{0.56 \text{ Ag}}{100 \text{ ore}} = 4480 \text{ g Ag} \rightarrow 4500 \text{ g Ag}$$

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 $12 \text{ in} = 1 \text{ ft}$ ,  $5280 \text{ ft} = 1 \text{ mile}$ ,  $3 \text{ feet} = 1 \text{ yd}$ ,  $60 \text{ sec} = 1 \text{ min}$ ,  $1 \text{ hr} = 60 \text{ min}$ ,  $4 \text{ quarts} = 1 \text{ gal}$ ,  $2 \text{ pints} = 1 \text{ quart}$

1. (6 Pts) Perform each of the following conversions. You must show the complete setup.

a. Convert 55 nL to mL.

$$\frac{55 \text{ nL}}{1} \times \frac{10^{-9} \text{ m}}{1 \text{ n}} \times \frac{1 \text{ mL}}{10^{-3} \text{ m}} = 55 \times 10^{-6} \text{ mL} \quad (5.5 \times 10^{-5} \text{ mL})$$

b. Convert 85 mL to  $\mu\text{L}$ .

$$\frac{85 \text{ mL}}{1} \times \frac{10^{-3} \text{ m}}{1 \text{ mL}} \times \frac{1 \mu\text{L}}{10^{-6} \text{ m}} = 85 \times 10^3 \mu\text{L} \quad (8.5 \times 10^4 \mu\text{L})$$

2. (4 Pts) Assume each of following numbers are measurements. Perform the indicated operations and then report the answer with the proper number of significant figures.

Limited to the 0.01 place

a.  $3.14 \text{ cm} + 12.526 \text{ cm} + 0.12 \text{ cm} = 15.786 \text{ cm} \quad (15.79 \text{ cm})$

3 s.f.

b.  $1.25 \text{ cm} \times 6.12 \text{ cm} \times 1.145 \text{ cm} = 8.75925 \text{ cm}^3 \quad (8.76 \text{ cm}^3)$

3. (5 Pts) A poster measures 33 inches by 45 inches. Determine its area in square cm.

$$\frac{33 \text{ in}}{1} \times \frac{2.54 \text{ cm}}{1 \text{ in}} \times \frac{45 \text{ in}}{1} \times \frac{2.54 \text{ cm}}{1 \text{ in}} = 9580 \text{ cm}^2 \quad (9600 \text{ cm}^2)$$

4. (5 Pts) How many inches are in 67 miles?

$$\frac{67 \text{ miles}}{1} \times \frac{5280 \text{ ft}}{1 \text{ mile}} \times \frac{12 \text{ in}}{1 \text{ ft}} = 4245120 \text{ in}$$

5. (5 Pts) A sample of silver ore was found to contain 0.066 % silver by mass. How many grams of silver can be recovered 500.0 kg of ore?

$$\frac{500.0 \text{ kg ore}}{1} \times \frac{10^3 \text{ g}}{1 \text{ kg}} \times \frac{0.066 \text{ Ag}}{100 \text{ ore}} = 330 \text{ g Ag}$$