	Q 9 CHM152 25 Pts fall 2005 ksp Name:
1.	(8 Pts) The solubility of magnesium phosphate is $2.27 \times 10^{-3}$ g/1.0 L of solution. What is the $K_{sp}$ for Mg <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> ?
2.	Lead(II) iodide, PbI <sub>2</sub> , is an ionic compound with a solubility product constant $K_{sp}$ of 7.9 × 10 <sup>-9</sup> . Calculate the molar solubility of this compound in a. (8 Pts) pure water.
	b. (8 Pts) 0.50 mol L <sup>-1</sup> KI solution.